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## An Activity Series Lab Answers

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Activity Series Lab Answers Introduction. The purpose of the lab was to find which metal is the most reactive and which metal is the least reactive. Materials. Procedure. Use a pipette to fill each of the four wells in column 1 of the reaction plate with 2mL of 1.0M aluminum... Results. ...

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element that is being replaced. Therefore, it is useful to have a list of elements in order of their relative reactivities. The activity series is a list of elements in decreasing order of their reactivity. Since metals replace other metals, while nonmetals replace other nonmetals, they each have a separate activity series.

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Play this game to review Chemical Reactions. Based on the activity series, will this reaction occur? Ni (s) + H 2 O (l)

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Simple Activity Series of Metals Lab for Standard Chemistry

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Activity Series of a Metal Lab - YouTube

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When an atom gains electrons, it is reduced. Metals higher on the activity series are more likely to react relative to those lower on the activity series. The activity series can be used to predict products of reactions, and to predict if a reaction will even occur. In this experiment, different metals were tested for their reactivity. It was recorded if a reaction occurred or not, so that an activity series could be created. Data & Results

Chemistry Lab Report (The Activity Series) – Sarah Jackson

Differences in metal reactivity can be represented as a metal activity series. • Determine whether a reaction will occur between a metal and a solution containing the ions of another metal, given a metal activity series containing both metal.

The Reactivity Series (solutions, examples, activities ...  
If the color of the halogen is produced there was no reaction. From this lab one can conclude that the activity series for the metals, from most active to least, is Magnesium (Mg), Zinc (Zn), Lead (Pb), Copper (Cu), Silver (Ag). The halogen activity series is Chlorine (Cl),

Bromine (Br), Iodine (I). No sources of error.

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For example:  $A + BX \rightarrow AX + B$  if A is above B in the activity series, but  $A + BX \rightarrow$  No reaction if A is below B in the activity series. Why iron displaces copper from copper sulphate...

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Virtual Lab: Activity Series - Mr. Rong's Chemistry

$2I^- (aq) + Cl_2 (aq) \rightarrow 2Cl^- (aq) + I_2 (aq)$   
 $2I^- (aq) + Br_2 (aq) \rightarrow 2Br^- (aq) + I_2 (aq)$   
8.  $Cl_2$  (most reactive),  $Br_2$ ,  $I_2$  (least reactive) is in the same order as in the activity series in the textbook. 9. The standard reduction potential table lists the halogens in the order  $Cl_2$ ,  $Br_2$ ,  $I_2$  making both lists the same.

The Activity Series. Single-replacement reactions only occur when the element that is doing the replacing is more reactive than the element that is being replaced. Therefore, it is useful to have a list of elements in order of their relative reactivities. The activity series is a list of elements in decreasing order of their reactivity. Since metals replace other metals, while nonmetals replace other nonmetals, they each have a separate activity series.

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