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## Chapter 9 Review Stoichiometry Section 2 Answers

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yield is 25 g. Calculate the percentage yield. 2. 6.0 mol of N<sub>2</sub> are mixed with 12.0 mol of H<sub>2</sub> according to the terms.

following equation:  $N_2(g) + 3H_2(g) \rightarrow 2NH_3(g)$

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The coefficients in a chemical equation represent the

(a) masses in grams of all reactants and products. (b)

relative number of moles of reactants and products.

(c) number of atoms of each element in each

compound in a reaction.

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following equation represents a laboratory preparation

for oxygen gas: CHAPTER 9 REVIEW fewer steps are

required to solve stoichiometry problems when. ...

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Stoichiometry. SECTION 2. PROBLEMS Write the

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$2KClO_3(s) \rightarrow 2KCl(s) + 3O_2(g)$  CHAPTER 9 REVIEW ...

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